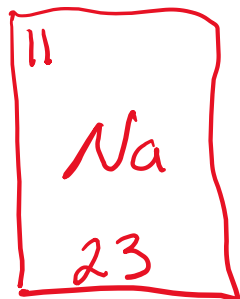


Bohr Models

- a 2D model/diagram of an atom that allows us to visualize the relative arrangement of sub-atomic particles in an atom
- by understanding the structure and position of subatomic particles (specifically the electrons) it helps us to understand, explain & predict an element's properties (reactivity) and how it will react with other elements (bonding ratio)
- the model shows how many electrons occupy each specific energy shell/level (orbit)

Rules for Bohr Models



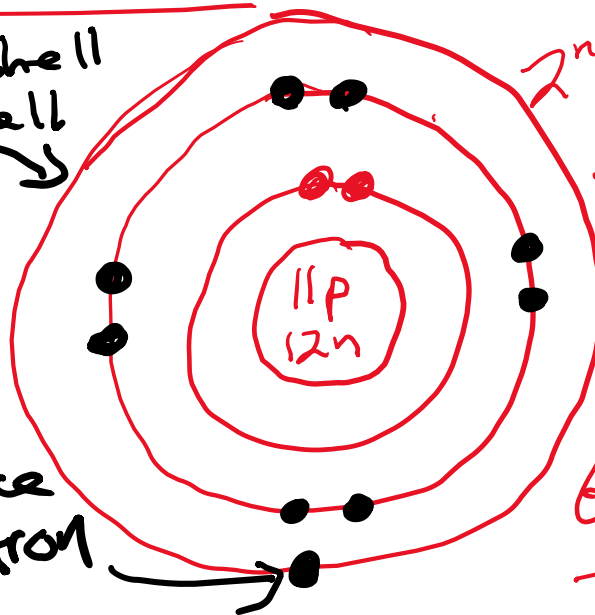
outermost shell
= valence shell

atom = 11e

$23 - 11 = 12$

11 p
12 n

valence
electron



1st shell = 2e

2nd shell = 8e

3rd shell = 8e

4th shell = 1e

5th shell = 18e

6th " = 32e

7th " = 32e

1. must fill innermost shell.
before adding next shell

2. use 4 points of compass
when placing electrons

3. # of p + n goes in nucleus

* row 2 has 8 elements
which = the # of electrons
in the second energy shell.